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composed of positive and negative ions
that are combined so that the numbers
of positive and negative charges are
equal. Formula Unit. the simplest
collection of atoms from which an ionic
compound's formula can be established.

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following questions in the space

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provided. c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) divided by the mass of 1 mol. nRT (c) multiplied by 22.4 L. (d) divided by 22.4 L. For the expression $V =$ (a) increasing P

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Modern Chemistry 9 Chemical Bonding
CHAPTER 6 STUDY GUIDE Chemical
Bonding SECTION 3 IONIC BONDING AND
IONIC COMPOUNDS SHORT ANSWER
Answer the following questions in the
space provided. 1. _____ The notation for
sodium chloride, NaCl, stands for one (a)
formula unit.

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Modern Chemistry Chapter 6 Review Answers Section 3

CHAPTER 3 REVIEW Atoms: The Building
Blocks of Matter SECTION 3 SHORT
ANSWER Answer the following questions
in the space provided. 1. Explain the
difference between the mass
number and the atomic number of a
nuclide. Mass number is the total

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number of protons and neutrons in the nucleus of an isotope.

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The Periodic Law - Jefferson Township

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a

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nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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